

1,4-Bis(methylsulfanyl)naphthalene

İsmail Çelik,^a Mehmet Akkurt,^{b*} Ayşegül Şenocak,^c
Osman Çakmak,^c Laura Torre-Fernández^d and Santiago
García-Granda^d

^aDepartment of Physics, Faculty of Arts and Sciences, Cumhuriyet University, 06532 Sivas, Turkey, ^bDepartment of Physics, Faculty of Arts and Sciences, Erciyes University, 38039 Kayseri, Turkey, ^cDepartment of Chemistry, Faculty of Arts and Sciences, Gaziosmanpaşa University, 60240 Tokat, Turkey, and ^dDepartamento Química Física y Analítica, Facultad de Química, Universidad Oviedo, C/ Julián Clavería, 8, 33006 Oviedo (Asturias), Spain
Correspondence e-mail: akkurt@erciyes.edu.tr

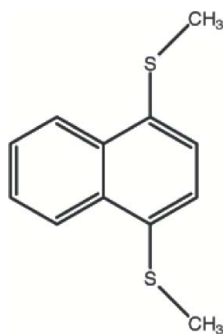
Received 28 April 2009; accepted 17 May 2009

Key indicators: single-crystal synchrotron study; $T = 153$ K; mean $\sigma(\text{C}-\text{C}) = 0.002$ Å; R factor = 0.036; wR factor = 0.109; data-to-parameter ratio = 23.7.

The molecule of the title compound, $\text{C}_{12}\text{H}_{12}\text{S}_2$, is close to planar, with the methyl C atoms deviating by 0.019 (1) and 0.221 (2) Å from the naphthalene mean plane. In the crystal structure, the shortest $\text{S} \cdots \text{S}$ contact of 3.6864 (9) Å is longer than the van der Waals contact distance.

Related literature

For general background, see: Underhill (1992); Öncü *et al.* (2006). For related structures, see: Noreland *et al.* (1992, 1993). For bond-length data, see: Allen *et al.* (1987). For van der Waals radii, see: Bondi (1964).



Experimental

Crystal data

 $\text{C}_{12}\text{H}_{12}\text{S}_2$
 $M_r = 220.34$

Monoclinic, $P2_1/c$
 $a = 15.203$ (3) Å
 $b = 10.246$ (2) Å
 $c = 7.1750$ (14) Å
 $\beta = 99.43$ (3)°
 $V = 1102.6$ (4) Å³

$Z = 4$
Synchrotron radiation
 $\lambda = 0.75140$ Å
 $\mu = 0.44$ mm⁻¹
 $T = 153$ K
 $0.11 \times 0.08 \times 0.02$ mm

Data collection

Bruker *P4* diffractometer
Absorption correction: part of the refinement model (ΔF)
(*XABS2*; Parkin *et al.*, 1995)
 $T_{\min} = 0.832$, $T_{\max} = 0.991$

5197 measured reflections
3013 independent reflections
2619 reflections with $I > 2\sigma(I)$
 $R_{\text{int}} = 0.019$

Refinement

$R[F^2 > 2\sigma(F^2)] = 0.036$
 $wR(F^2) = 0.109$
 $S = 1.07$
3013 reflections

127 parameters
H-atom parameters constrained
 $\Delta\rho_{\text{max}} = 0.25$ e Å⁻³
 $\Delta\rho_{\text{min}} = -0.25$ e Å⁻³

Data collection: *XSCANS* (Bruker, 1996); cell refinement: *SCALEPACK* (Otwinowski & Minor, 1997); data reduction: *DENZO* (Otwinowski & Minor, 1997) and *SCALEPACK*; program(s) used to solve structure: direct methods using *SIR2004* (Burla *et al.*, 2005); program(s) used to refine structure: *SHELXL97* (Sheldrick, 2008); molecular graphics: *ORTEP-3 for Windows* (Farrugia, 1997); software used to prepare material for publication: *WinGX* (Farrugia, 1999).

SGG gratefully acknowledges financial support from the MEC, projects MAT2006-01997 and Consolider Ingenio-2010, 'Factoría Española de Cristalización', for the latter of which beam time provided by BM16 (ESRF) under proposal 16-01-706 is gratefully acknowledged.

Supplementary data and figures for this paper are available from the IUCr electronic archives (Reference: HB2962).

References

- Allen, F. H., Kennard, O., Watson, D. G., Brammer, L., Orpen, A. G. & Taylor, R. (1987). *J. Chem. Soc. Perkin Trans. 2*, pp. S1–19.
- Bondi, A. (1964). *J. Phys. Chem.* **68**, 441–451.
- Bruker (1996). *XSCANS*. Bruker AXS Inc., Karlsruhe, Germany.
- Burla, M. C., Caliandro, R., Camalli, M., Carrozzini, B., Cascarano, G. L., De Caro, L., Giacovazzo, C., Polidori, G. & Spagna, R. (2005). *J. Appl. Cryst.* **38**, 381–388.
- Farrugia, L. J. (1997). *J. Appl. Cryst.* **30**, 565.
- Farrugia, L. J. (1999). *J. Appl. Cryst.* **32**, 837–838.
- Noreland, J., Olovsson, G. & Olovsson, I. (1992). *Acta Cryst.* **C48**, 1459–1462.
- Noreland, J., Olovsson, G. & Olovsson, I. (1993). *Acta Cryst.* **C49**, 168–171.
- Öncü, A., Büyükkıdan, N. S., Büyükkıdan, B. & Çakmak, O. (2006). *Turk. J. Chem.* **30**, 235–241.
- Otwinowski, Z. & Minor, W. (1997). *Methods in Enzymology*, Vol. 276, *Macromolecular Crystallography*, Part A, edited by C. W. Carter Jr & R. M. Sweet, pp. 307–326. New York: Academic Press.
- Parkin, S., Moezzi, B. & Hope, H. (1995). *J. Appl. Cryst.* **28**, 53–56.
- Sheldrick, G. M. (2008). *Acta Cryst.* **A64**, 112–122.
- Underhill, A. E. (1992). *J. Mater. Chem.* **2**, 1–11.

supplementary materials

Acta Cryst. (2009). E65, o1376 [doi:10.1107/S1600536809018650]

1,4-Bis(methylsulfanyl)naphthalene

İ. Çelik, M. Akkurt, A. Senocak, O. Çakmak, L. Torre-Fernández and S. García-Granda

Comment

Molecular electronics have attracted considerable attention because they are used for production electronic switches, memory cells or sensors. The most extensive range of the set compounds is based on planar organic donor molecules containing S or Se atoms (*e.g.* Underhill, 1992; Noreland *et al.*, 1992; Noreland *et al.*, 1993; Öncü *et al.*, 2006. In this study, the synthesis and structure of the title compound, (I), are reported.

As shown in Fig. 1, the naphthalene group (C1–C10) of the title molecule (I) is essentially planar, with maximum deviations of -0.014 (1), -0.014 (1) and 0.018 (1) Å for C4, C8 and C6, respectively. Deviations from planarity in the title molecule which were calculated to the least-squares plane of the naphthalene group are: 0.014 (1) Å for S1, 0.112 (1) Å for S2, -0.221 (2) Å for C11 and -0.019 (1) Å for C12. The bond lengths and angles in (I) show normal values (Allen *et al.*, 1987). The molecular packing of (I) as seen down the *a* axis is illustrated in Fig. 2. There is only one S··S contact near the van der Waals contact distance, 3.6 Å (Bondi, 1964): S2··S2ⁱ 3.6864 (9) Å [symmetry code: (i) -*x* + 1, -*y*, -*z*].



Experimental

The complete reaction was carried out under nitrogen atmosphere and 195 K. *t*-BuLi (10 ml, 14 mmol) solution was added to 1,4-dibromonaphthalene (1 g, 3.5 mmol) solved in dry THF (12 ml) with continuous stirring. Meanwhile, red colour was observed because of LiBr formation. Later, to the resulting reaction mixture was added dropwise (CH₃S)₂ (1.24 ml, 14 mmol). After 1 h, the reaction was ended. Reaction mixture warming to room temperature was extracted with Et₂O (3 × 25 ml). The extract was dried over Na₂SO₄ and concentrated. 1,4-Bis(methylthio)naphthalene as white needle-like crystals was obtained by crystallization from crude product solved in dichloromethane:hexan (1:19) system (708 mg, 92%), m.p. 369–371 K (lit. 370–371 K).

Refinement

H atoms were positioned geometrically and allowed to ride on their parent atoms, with C—H = 0.93 Å and $U_{\text{iso}}(\text{H}) = 1.2U_{\text{eq}}(\text{C})$ for aromatic H atoms, and 0.9 Å and $U_{\text{iso}}(\text{H}) = 1.5U_{\text{eq}}(\text{C})$ for methyl H atoms.

Figures

-  Fig. 1. The molecular structure of (I), with displacement ellipsoids drawn at the 50% probability level.
-  Fig. 2. View of the molecular packing diagram of compound (I) viewed down *b* axis. H atoms have been omitted for clarity.

1,4-Bis(methylsulfanyl)naphthalene

Crystal data

$C_{12}H_{12}S_2$	$F_{000} = 464$
$M_r = 220.34$	$D_x = 1.327 \text{ Mg m}^{-3}$
Monoclinic, $P2_1/c$	Synchrotron radiation
Hall symbol: -P 2ybc	$\lambda = 0.75140 \text{ \AA}$
$a = 15.203 (3) \text{ \AA}$	Cell parameters from 40396 reflections
$b = 10.246 (2) \text{ \AA}$	$\theta = 1.0\text{--}27.0^\circ$
$c = 7.1750 (14) \text{ \AA}$	$\mu = 0.44 \text{ mm}^{-1}$
$\beta = 99.43 (3)^\circ$	$T = 153 \text{ K}$
$V = 1102.6 (4) \text{ \AA}^3$	Prism, colourless
$Z = 4$	$0.11 \times 0.08 \times 0.02 \text{ mm}$

Data collection

Bruker P4 diffractometer	3013 independent reflections
Radiation source: sealed tube	2619 reflections with $I > 2\sigma(I)$
Monochromator: graphite	$R_{\text{int}} = 0.019$
$T = 153 \text{ K}$	$\theta_{\text{max}} = 33.8^\circ$
ω scans	$\theta_{\text{min}} = 2.6^\circ$
Absorption correction: part of the refinement model (ΔF) (XABS2; Parkin <i>et al.</i> , 1995)	$h = -20 \rightarrow 20$
$T_{\text{min}} = 0.832$, $T_{\text{max}} = 0.991$	$k = -14 \rightarrow 14$
5197 measured reflections	$l = -9 \rightarrow 0$

Refinement

Refinement on F^2	Secondary atom site location: difference Fourier map
Least-squares matrix: full	Hydrogen site location: inferred from neighbouring sites
$R[F^2 > 2\sigma(F^2)] = 0.036$	H-atom parameters constrained
$wR(F^2) = 0.109$	$w = 1/[\Sigma^2(F_o^2) + (0.0553P)^2 + 0.1466P]$
$S = 1.07$	where $P = (F_o^2 + 2F_c^2)/3$
3013 reflections	$(\Delta/\sigma)_{\text{max}} = 0.001$
127 parameters	$\Delta\rho_{\text{max}} = 0.25 \text{ e \AA}^{-3}$
Primary atom site location: structure-invariant direct methods	$\Delta\rho_{\text{min}} = -0.25 \text{ e \AA}^{-3}$
	Extinction correction: none

Special details

Experimental. Cubic fit to $\sin(\theta)/\lambda$ - 24 parameters

Geometry. Bond distances, angles *etc.* have been calculated using the rounded fractional coordinates. All su's are estimated from the variances of the (full) variance-covariance matrix. The cell e.s.d.'s are taken into account in the estimation of distances, angles and torsion angles

Refinement. Refinement on F^2 for ALL reflections except those flagged by the user for potential systematic errors. Weighted R -factors wR and all goodnesses of fit S are based on F^2 , conventional R -factors R are based on F , with F set to zero for negative F^2 . The observed criterion of $F^2 > \sigma(F^2)$ is used only for calculating $-R$ -factor-obs *etc.* and is not relevant to the choice of reflections for refinement. R -factors based on F^2 are statistically about twice as large as those based on F , and R -factors based on ALL data will be even larger.

Fractional atomic coordinates and isotropic or equivalent isotropic displacement parameters (\AA^2)

	<i>x</i>	<i>y</i>	<i>z</i>	$U_{\text{iso}}^*/U_{\text{eq}}$
S1	0.90863 (2)	-0.04680 (4)	0.76728 (6)	0.0671 (2)
S2	0.53023 (2)	-0.06992 (3)	0.23860 (5)	0.0558 (1)
C1	0.85913 (9)	0.07840 (13)	0.3828 (2)	0.0578 (4)
C2	0.84525 (11)	0.14096 (16)	0.2131 (2)	0.0700 (5)
C3	0.76299 (11)	0.13355 (17)	0.0952 (2)	0.0714 (5)
C4	0.69472 (10)	0.06351 (14)	0.1496 (2)	0.0592 (4)
C5	0.70573 (8)	-0.00085 (11)	0.32680 (17)	0.0468 (3)
C6	0.63478 (8)	-0.07293 (11)	0.38797 (18)	0.0469 (3)
C7	0.65018 (9)	-0.13631 (14)	0.55763 (19)	0.0562 (4)
C8	0.73388 (9)	-0.13004 (14)	0.67599 (19)	0.0585 (4)
C9	0.80296 (8)	-0.05983 (12)	0.62517 (19)	0.0511 (4)
C10	0.79008 (8)	0.00626 (11)	0.44685 (18)	0.0478 (3)
C11	0.90238 (12)	-0.15800 (18)	0.9562 (3)	0.0818 (6)
C12	0.46292 (9)	-0.17829 (14)	0.3515 (2)	0.0625 (4)
H1	0.91490	0.08300	0.45830	0.0690*
H2	0.89120	0.18910	0.17560	0.0840*
H3	0.75440	0.17620	-0.02070	0.0860*
H4	0.64030	0.05810	0.06910	0.0710*
H7	0.60450	-0.18460	0.59600	0.0670*
H8	0.74230	-0.17440	0.79060	0.0700*
H11A	0.95780	-0.15720	1.04240	0.1230*
H11B	0.89100	-0.24440	0.90620	0.1230*
H11C	0.85490	-0.13220	1.02170	0.1230*
H12A	0.40410	-0.18270	0.27870	0.0940*
H12B	0.45940	-0.14650	0.47600	0.0940*
H12C	0.48920	-0.26370	0.36040	0.0940*

Atomic displacement parameters (\AA^2)

	U^{11}	U^{22}	U^{33}	U^{12}	U^{13}	U^{23}
S1	0.0534 (2)	0.0744 (3)	0.0682 (3)	-0.0142 (2)	-0.0060 (2)	0.0043 (2)

supplementary materials

S2	0.0519 (2)	0.0567 (2)	0.0553 (2)	-0.0028 (1)	-0.0020 (1)	0.0051 (1)
C1	0.0503 (7)	0.0558 (7)	0.0691 (8)	-0.0030 (5)	0.0149 (6)	0.0008 (6)
C2	0.0643 (8)	0.0693 (9)	0.0818 (10)	-0.0055 (7)	0.0280 (7)	0.0134 (7)
C3	0.0707 (9)	0.0773 (10)	0.0696 (9)	0.0035 (7)	0.0212 (7)	0.0236 (8)
C4	0.0578 (7)	0.0612 (7)	0.0593 (8)	0.0062 (5)	0.0117 (6)	0.0111 (6)
C5	0.0494 (6)	0.0417 (5)	0.0504 (6)	0.0037 (4)	0.0116 (4)	0.0001 (4)
C6	0.0453 (6)	0.0447 (5)	0.0497 (6)	0.0001 (4)	0.0047 (4)	-0.0012 (4)
C7	0.0490 (6)	0.0610 (7)	0.0571 (7)	-0.0094 (5)	0.0043 (5)	0.0098 (6)
C8	0.0540 (7)	0.0660 (8)	0.0528 (7)	-0.0088 (6)	0.0007 (5)	0.0117 (6)
C9	0.0473 (6)	0.0505 (6)	0.0537 (7)	-0.0036 (5)	0.0031 (5)	-0.0019 (5)
C10	0.0484 (6)	0.0411 (5)	0.0550 (6)	0.0002 (4)	0.0115 (5)	-0.0026 (5)
C11	0.0770 (10)	0.0848 (11)	0.0728 (10)	-0.0151 (9)	-0.0198 (8)	0.0164 (8)
C12	0.0536 (7)	0.0643 (8)	0.0667 (8)	-0.0097 (6)	0.0017 (6)	0.0023 (6)

Geometric parameters (Å, °)

S1—C9	1.7611 (14)	C9—C10	1.4324 (18)
S1—C11	1.785 (2)	C1—H1	0.9300
S2—C6	1.7657 (13)	C2—H2	0.9300
S2—C12	1.7911 (15)	C3—H3	0.9300
C1—C2	1.362 (2)	C4—H4	0.9300
C1—C10	1.4204 (19)	C7—H7	0.9300
C2—C3	1.392 (2)	C8—H8	0.9300
C3—C4	1.370 (2)	C11—H11A	0.9600
C4—C5	1.4176 (19)	C11—H11B	0.9600
C5—C6	1.4341 (17)	C11—H11C	0.9600
C5—C10	1.4247 (18)	C12—H12A	0.9600
C6—C7	1.3656 (19)	C12—H12B	0.9600
C7—C8	1.411 (2)	C12—H12C	0.9600
C8—C9	1.3707 (19)		
S2...S2 ⁱ	3.6864 (9)		
C9—S1—C11	103.55 (8)	C1—C2—H2	120.00
C6—S2—C12	103.82 (7)	C3—C2—H2	120.00
C2—C1—C10	121.32 (14)	C2—C3—H3	120.00
C1—C2—C3	120.65 (15)	C4—C3—H3	120.00
C2—C3—C4	120.16 (14)	C3—C4—H4	120.00
C3—C4—C5	121.03 (14)	C5—C4—H4	119.00
C4—C5—C6	121.99 (12)	C6—C7—H7	119.00
C4—C5—C10	118.74 (12)	C8—C7—H7	119.00
C6—C5—C10	119.27 (11)	C7—C8—H8	119.00
S2—C6—C5	116.91 (10)	C9—C8—H8	119.00
S2—C6—C7	123.72 (10)	S1—C11—H11A	109.00
C5—C6—C7	119.34 (12)	S1—C11—H11B	109.00
C6—C7—C8	121.39 (13)	S1—C11—H11C	109.00
C7—C8—C9	121.24 (13)	H11A—C11—H11B	110.00
S1—C9—C8	123.49 (11)	H11A—C11—H11C	109.00
S1—C9—C10	117.36 (10)	H11B—C11—H11C	109.00
C8—C9—C10	119.15 (12)	S2—C12—H12A	109.00

C1—C10—C5	118.08 (12)	S2—C12—H12B	109.00
C1—C10—C9	122.34 (12)	S2—C12—H12C	109.00
C5—C10—C9	119.58 (11)	H12A—C12—H12B	109.00
C2—C1—H1	119.00	H12A—C12—H12C	109.00
C10—C1—H1	119.00	H12B—C12—H12C	109.00
C11—S1—C9—C8	7.08 (14)	C10—C5—C6—C7	-1.81 (18)
C11—S1—C9—C10	-172.38 (11)	C4—C5—C10—C1	0.28 (18)
C12—S2—C6—C5	175.91 (10)	C4—C5—C10—C9	-179.02 (12)
C12—S2—C6—C7	-5.94 (13)	C6—C5—C10—C1	-179.94 (12)
C10—C1—C2—C3	-1.4 (2)	C6—C5—C10—C9	0.76 (17)
C2—C1—C10—C5	1.0 (2)	S2—C6—C7—C8	-176.70 (11)
C2—C1—C10—C9	-179.70 (14)	C5—C6—C7—C8	1.4 (2)
C1—C2—C3—C4	0.4 (2)	C6—C7—C8—C9	0.1 (2)
C2—C3—C4—C5	0.9 (2)	C7—C8—C9—S1	179.38 (11)
C3—C4—C5—C6	178.97 (14)	C7—C8—C9—C10	-1.2 (2)
C3—C4—C5—C10	-1.3 (2)	S1—C9—C10—C1	0.93 (17)
C4—C5—C6—S2	-3.80 (16)	S1—C9—C10—C5	-179.80 (10)
C4—C5—C6—C7	177.96 (13)	C8—C9—C10—C1	-178.56 (13)
C10—C5—C6—S2	176.43 (9)	C8—C9—C10—C5	0.71 (18)

Symmetry codes: (i) $-x+1, -y, -z$.

Hydrogen-bond geometry ($\text{\AA}, ^\circ$)

<i>D</i> —H \cdots <i>A</i>	<i>D</i> —H	H \cdots <i>A</i>	<i>D</i> \cdots <i>A</i>	<i>D</i> —H \cdots <i>A</i>
C1—H1 \cdots S1	0.93	2.60	3.0235 (16)	108
C4—H4 \cdots S2	0.93	2.58	3.0090 (17)	108

Fig. 1

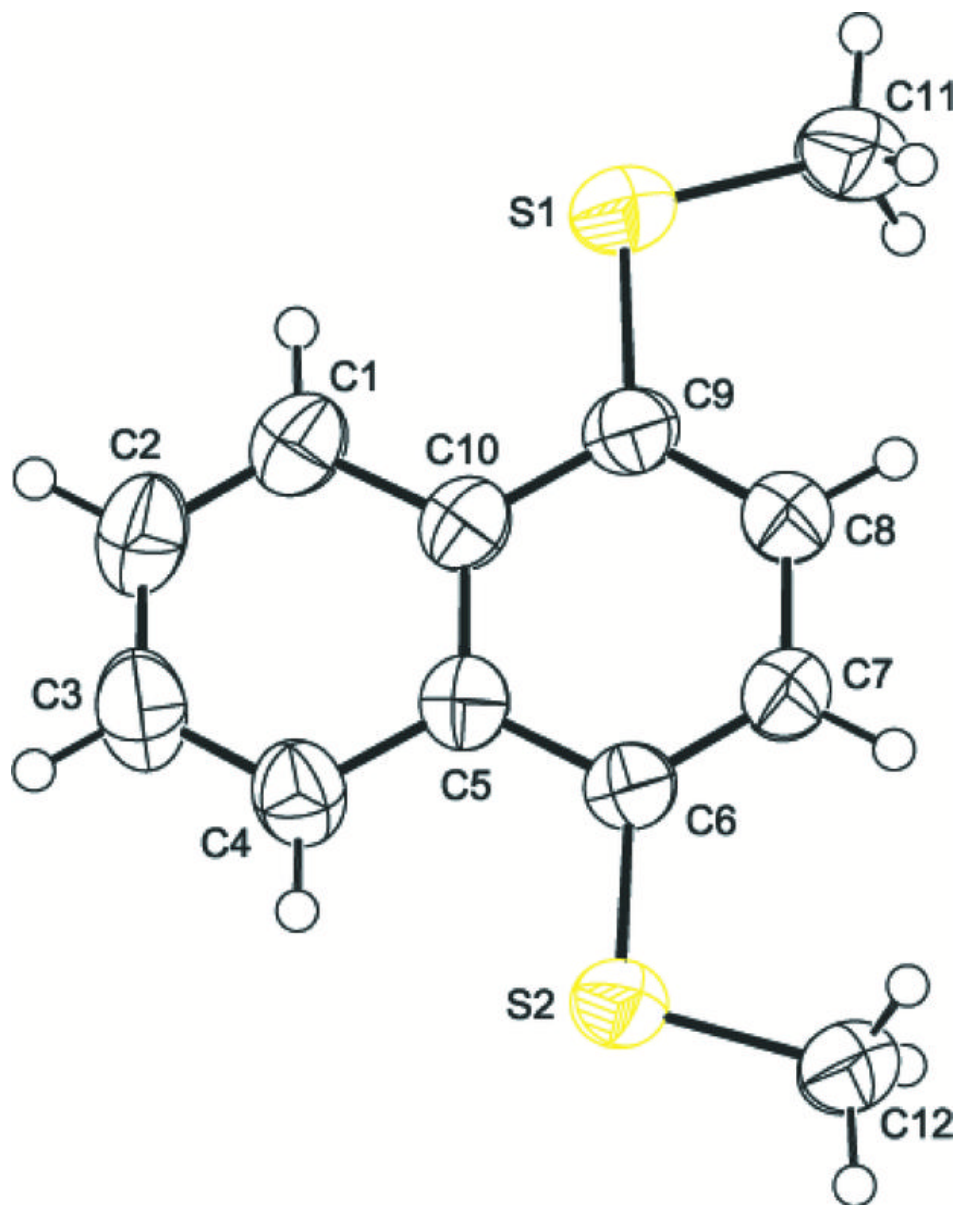


Fig. 2

